

### **chapter 10 chemical quantities pdf**

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### **CHAPTER 10: Chemical Quantities**

Chemical Quantities THE MOLE AND QUANTIFYING MATTER 10.1 The Mole: A ... Use the chemical formula to find the number of atoms in one molecule and multiply ... The quantities 1 mol and 22.4 L can be used in conversion factors that change moles to

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Balanced chemical equations tell us in what proportions on a mole basis substances combine; since the molar masses of C (s) and O<sub>2</sub> (g) are different, 1 g of O<sub>2</sub> could not represent the same number of moles as 1 g of C.

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Chapter 10: Chemical Quantities Section 10.1: The mole, a measurement of matter. How do you measure the amount of matter One way is to simply count the number of ... A mole (mol) of a substance is  $6.02 \times 10^{23}$  representative particles of that substance and is the SI unit for measuring the amount of a substance.

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Chapter 10 Chemical Quantities. 259. 33. What is the empirical formula of a compound that is 27.3% C and 72.7% O? 34. A compound consisting of 56.38% phosphorus and 43.62% oxygen has a molar mass of 219.9 g/mol. Determine its molecular formula. D. Essay. Write a short essay for the following. 35.

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