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Electron Energy and Light 5 20. Label the remaining drawings in Model 3 with the electron transitions that are occurring ($n=?$ to $n=?$), the wavelengths and corresponding colors as given in example A in Model 3.

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This electron moves from a (lower/higher) energy state to a (lower/higher) energy state. b. Inc. Identify the drawing in Model 3 that depicts a hydrogen atom with an electron moving from energy level 5 to energy level 2.

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(b) This electron moves from a lower / higher energy state to a lower / higher energy state. (c) Light is absorbed / released in the 3 to 2 electron transition. 23.

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Electron Energy and Light . Model 3 - 543 Bohr Model of a Hydrogen Atom 2 19. Identify the drawing in Model 3 that depicts a hydrogen atom with an electron moving from energy level 5 to energy level 2. Refer to Models 1 and 2 for the following questions. a. Label the picture with to $n=2$ " and list the corresponding color of light emitted.

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LIGHT QUESTIONS and ANSWERS 1. What does it mean to say an atom is excited? An electron or electrons have more energy than their ground state, which is where they would

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The energy level of a bound electron determines the orbital it occupies, and the color reflects the probability of finding the electron at a given position. An electron can be bound to the nucleus of an atom by the attractive Coulomb force.

Electron - Wikipedia

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