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For non-volatile solutes :  $P(\text{solution}) = x \text{ solvent } P_o$  or relative lowering of vapor pressure,  $(P_o - P)/P_o = x_{\text{solute}}$ . A more useful form is  $(P_o - P)/P = n/N$  where  $n$  = total number of moles of all the free solute species in the solution finally (i.e. at equilibrium). Three cases arise. (i) Non-electrolyte is dissolved e.g. glucose or urea.

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Non-aqueous solutions: The solution in which solvent other than water is used is called as non-aqueous solution. Concentration of solutions: It is defined as the amount of solute dissolved in specific amount of solvent. Dilute solutions: The solutions containing relatively less amount of solute are called as dilute solutions.

### **Solutions and Colligative Properties - Saptarshi Classes**

A 1-molal solution contains one mole of solute per 1 kg of solvent. Molality is a hybrid concentration unit, retaining the convenience of mole measure for the solute, but expressing it in

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